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2.17 - 2.21 The solutions for these problems are the solutions for problems 2.15 - 2.19 in the 2nd edition Solutions Manual. 2.22 From Table 2.3, $K_{sp, CaSO_4}$ is 2×10^{-5} and $K_{sp} = [Ca^{2+}] [SO_4^{2-}] = [SO_4^{2-}]^2$ At saturation: $[SO_4^{2-}] = K_{sp}^{1/2} = (2 \times 10^{-5})^{1/2} = 0.00447 \text{ mol/L}$ Check if saturation is exceeded: $\frac{0.00368 \text{ g}}{136 \text{ mole L}} \times 5.0 = 0.00134 \text{ mol/L}$ Since saturation is not exceeded the sulfate concentration is 0.00368 mol/L

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